

# Storage time of all-vanadium liquid energy storage power station

What materials are used to make vanadium redox flow batteries?

Image: CellCube. Samantha McGahan of Australian Vanadium writes about the liquid electrolyte which is the single most important material for making vanadium flow batteries, a leading contender for providing several hours of storage, cost-effectively. Vanadium redox flow batteries (VRFBs) provide long-duration energy storage.

How many litres of vanadium can be produced a year?

Primary vanadium producer Bushveld Minerals in South Africa is completing construction of its BELCO electrolyte plant which is expected to start operation in H1 2023, with an initial capacity of eight million litres per year. This production can be expanded to deliver 32 million litres per year.

Is vanadium a sustainable solution?

US Vanadium can recycle spent electrolyte from VRFBs at a 97% vanadium recovery rate. This makes the VRFB a truly sustainable solution- the vanadium resource is only being borrowed from future generations, not consumed at its expense. One of the main costs affecting vanadium electrolyte is the price of moving it.

How many litres of vanadium is a MWh?

One megawatt-hour (1MWh) of stored energy equals approximately 68,000 litres of vanadium electrolyte or 9.89 tonnes of vanadium pentoxide ( $V_2O_5$ ), which can include a proportion of vanadium (III) oxide ( $V_2O_3$ ) depending on whether a chemical or electrical method of production is used.

Which equation is used to study vanadium ion concentration variations?

Combining the energy conservation equation, Nernst equation, mass balance equation and Bernoulli equation, a numerical ordinary differential equation (ODE) can be established to study the vanadium ion concentration variations with flow rate, electrical characteristics, hydraulic system design and electrolyte temperature.

What is an all-vanadium flow battery (VFB)?

The all-vanadium flow battery (VFB) employs  $V^{2+} / V^{3+}$  and  $VO^{2+} / VO^{3+}$  redox couples in dilute sulphuric acid for the negative and positive half-cells respectively. It was first proposed and demonstrated by Skyllas-Kazacos and co-workers from the University of New South Wales (UNSW) in the early 1980s.



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